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Concepts: Terms in this

set (57) Dalton. The

atom is a tiny,

indestructible particle

with no internal

structure. Thomson.

The atom is a sphere of

positive electrical

charge with electrons

embedded in the

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sphere.

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CHAPTER 5 What You'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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Chapter 5 - Electrons
in Atoms. Chapter 5 -
Electrons in Atoms.

Jennie L. Borders.

Section 5.1 - Models of
the Atom. The

Rutherford's model of
the atom did not

explain how an atom
can emit light or the
chemical properties of

an atom. Plum Pudding

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Model Rutherford's
Model. The Bohr Model.

Chapter 5 - Electrons in Atoms

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:

Aufbau principle - electrons enter the lowest energy first.

This causes difficulties because of the overlap of orbitals of different energies - follow the

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diagram! Pauli

Exclusion Principle - at

most 2 electrons per
orbital ...

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Chapter 5 - electrons in
atoms (handouts)

Chapter 6 - periodic
table & trends

(handouts) Chapters

7/9 - ionic bonding &
naming (handouts)

Chapters 8/9 - covalent
bonding & chemical

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names & formulas
(handouts) Chapters

8/15 - VSEPR/polar
bonding/IMFs

(handouts) Chapter 10
- moles (handouts)

Chapter 11 - reactions
(handouts) Chapter 12

...

Science / Chapter 5 - electrons in atoms (handouts)

Chapter 5: Electrons in
Atoms Models of the
Atom Rutherford used
existing ideas about

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the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Cristina Chemistry I H.

Terms in this set (82)

Rutherford's model-

also called nuclear

model-did not explain

how the electrons were

arranged

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in Atoms. Chapter 5 -

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Section 5.1 – Models of
the Atom. The
Rutherford's model of
the atom did not
explain how an atom
can emit light or the
chemical properties of
an atom. Plum Pudding
Model Rutherford's
Model. The Bohr Model.

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Electrons in Atoms

Section 55.1.1 Figure

5.1 Different elements
can have similar
reactions with water.

Objectives Compare
the wave and particle
natures of light. Define
a quantum of energy,
and explain how it is
related to an energy
change of matter.

Contrast continuous

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electromagnetic

**Chapter 5: Electrons
in Atoms**

Chapter 5 Electrons in
Atoms. Educators. AY
IB Chapter Questions.

02:25. Problem 1

Objects get their colors
from reflecting only
certain wavelengths
when hit with white
light. Light reflected
from a green leaf is
found to have a
wavelength of 4.90

$\times 10^{-7}$

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\mathrm{m} . \$ What
is the frequency of the
light? ...

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How are electrons
arranged in an atom?
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times. 0. 10th grade .

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Chapter 5“Electrons in
Atoms ... Describe the
energies and positions
of electrons according
to the quantum
mechanical model.

Section 5.1Models of
the Atom. OBJECTIVES:

Describe how the

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shapes of orbitals related to different sublevels differ. Ernest Rutherford's Model.

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can each p orbital

hold? Chapter 5:

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